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(54) Title: SURFACE CLEANER

(57) Abstract: The present invention relates to photocatalytic compositions effective to degrade soils deposited on a surface, to reduce the accumulation of soils on a surface, and to act as an antimicrobial agent. The invention also describes methods employing said compositions, and uses thereof.

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SURFACE CLEANER

The present invention relates to sensitising agents and compositions effective to  
5 degrade soils deposited on a surface, methods employing said agents and compositions,  
and uses thereof.

Cleaning compositions intended for general and specific uses are well known in  
the art. Such compositions will normally comprise one or more surfactants, solvents  
10 thickening agents, abrasive particles, bleaching agents, disinfectant/antibacterial agents,  
perfumes, preservatives and colouring agents. Although these compositions are effective  
at removing soils, inevitably resoiling occurs after cleaning and thus recleaning is  
required.

15 A means to reduce the frequency of cleaning and recleaning would thus be  
advantageous. In addition it would be beneficial if one could reduce the rate of  
accumulation of surface soils in the first instance. The present invention seeks to address  
these problems.

20 The photocatalysed degradation of organic environmental pollutants in the  
presence of a semiconductor such as titanium dioxide or zinc oxide is well known (Ollis  
et al, Environ. Sci. and Technol., 12 (1991) 1522; Heller Am. Chem. Res. 28 (1995),  
503). However, the chemical characteristics of these semiconductors necessitate  
excitation of these metals in the ultraviolet region of the spectrum in order for the  
25 degradation of the pollutants to occur. This requirement therefore makes the use of  
photocatalysed degradation of soils on surfaces within a residential environment both  
potentially hazardous and impractical.

The present inventors have found, however, that the use of a sensitising agent in  
30 addition to the light absorbing material reduces the amount of energy required to be  
absorbed by said light absorber in order for charge separation to take place, and  
subsequently for the photocatalysed degradation of surface soils to occur. The present  
inventors have found that ambient light, for example sunlight or artificial light is

sufficient in the presence of a sensitising agent and a light absorbing material to induce such a degradation.

The present inventors have found, in addition, that the use of highly conjugated heterocyclic complexes such as polypyridine, macrocycle or phthalocyanines with various centrally coordinated atoms such as Ru, Fe and Si can be used to sensitise a light absorbing agent (such as titanium dioxide or zinc oxide) not only when the light absorbing agent is coated onto a surface, but also when the agent is in solution. This makes the use of a light absorbing agent in conjunction with certain metal complexes in solution ideally suited for applying to a surface to provide a residue which will photocatalyse the decomposition of surface soils, and will also reduce the rate of accumulation of soils.

Previously, the photosensitisation of nanocrystalline titanium dioxide (TiO<sub>2</sub>) films by polyimide bearing pendant substituted-Ru (bpy)<sub>3</sub> +2 groups has been reported (Osora et al, J. Photochem. Photobiol. B: Biol. 43 (1998) 232). In this study it was found that these photosensitised complexes could degrade methylene blue.

Additionally, the use of metal complexes as photosensitisers in electrochemical cells is well established (Kalyanasundaram et al. Photo. Sens. And Photocat Using Inorg. And Organomet. Comps, 247-271). Now the use of particulate semiconductors such as TiO<sub>2</sub> with sensitisers is resulting in the development of a new class of solar cell (Graetzel et al, Nature, 1991, v353, 737).

Graetzel et al (JACS V107, (1985), 2988 showed that tris(2,2'-bipyridyl - 4,4'-dicarboxylate) ruthenium (II) dichloride is a superior sensitising agent for charge injection into titania at acid pH compared to tris(2,2'-bipyridyl) ruthenium (II) dichloride due to the former having carboxylate anions capable of binding with titania under acid conditions. In this same work it was also shown that the sensitising properties of tris(2,2'-bipyridyl) ruthenium (II) dichloride improve at pH 7 compared with lower pH.

Bendig et al (J Photochem Photobiology A: Chemistry 108 (1997) 89), describe the sensitised photocatalytic oxidation of herbicides using tris(2,2'-bipyridyl - 4,4'-dicarboxylate) ruthenium (II) dichloride, tris(2,2'-bipyridyl) ruthenium (II) dichloride and

a methylated form of the latter. Bendig et al showed that only tris(2,2'-bipyridyl - 4,4'-dicarboxylate) ruthenium (II) dichloride is active under acid conditions (pH 3) under their experimental conditions, and the authors presented data showing that tris(2,2'-bipyridyl - 4,4'-dicarboxylate) ruthenium (II) dichloride desorbs progressively as the pH is raised whereas tris(2,2'-bipyridyl) ruthenium (II) dichloride is more strongly absorbed with increasing pH. The authors explain the results on the basis that titania has a point of zero charge (PZC) such that the surface is positively charged below approximately pH6 and negatively charged at higher pH. Accordingly, under acid conditions sensitising agents carrying a negatively charged group can bind via electrostatic interaction, whereas positively charged groups will tend to be repelled. Conversely, at pH greater than the PZC value for titania, molecular moieties with positively charged groups will tend to bind more strongly with the TiO<sub>2</sub> surface.

It follows therefore, that for sensitisation to be most effective at a particular working pH, on semiconductors such as titania, zinc oxide, tin oxide etc, charged groups of the appropriate sign should be present on the absorbing sensitizer-molecule to promote binding. Thus for application at pH conditions where the semiconductor material has an excess negative charge, a sensitising molecule should preferably have a positively charged group or groups in its structure.

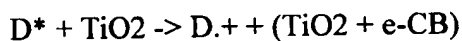
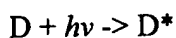
In the first aspect, the present invention provides a composition comprising a photocatalyst and a metal complex sensitizer comprising a ligand with a conjugated  $\pi$  system which absorbs light substantially in the visible and/or the infrared region of the spectrum, effective to deposit a functional residue of said composition on a surface.

The term 'functional residue', in the context of the present invention means a residue or layer of photocatalytic composition provided on a surface whereby soils deposited on the residue or layer or soils which are present on the surface prior to the deposition of the residue or layer are subject to a photocatalytic or other photochemical oxidation, reduction, free radical or other photochemical reaction effective to substantially break down, or otherwise decompose the soil. In effect, the cleaning process continues after the conventional act of soil removal is completed. In addition, these reactions may also provide an ongoing antibacterial effect that continues after the physical cleaning process has been completed. Finally, if a functional residue of

photocatalytic material is applied to a substantially clean or sterile surface then the rate of accumulation of soils on the surface will be reduced.

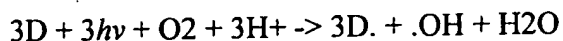
5 The term photocatalytic agent in the context of the present invention refers to an agent that has a favourable combination of electronic structure, light absorption properties, charge transport characteristics and excited-state lifetimes. Primary light absorbers for photocatalysis include but are not limited to semiconductor materials.

10 One model of dye sensitisation of the semi-conductor titanium dioxide, suggests that surface adsorbed dye molecules (sensitising agents) absorb visible light and inject electrons into the conduction band thus:



15

The conduction band electrons may then reduce oxygen to reactive species such as .OH radicals, which can rapidly attack organic molecules, i.e.



20

on  $\text{TiO}_2$

Alternatively or simultaneously  $D.^+$  may oxidise organic molecules. In this invention it will be understood by those skilled in the art that the sensitising agent is working in a catalytic manner i.e it is not significantly altered itself during the photocatalytic cleaning process, and is therefore active over a long period of time.

25

Suitable photocatalytic agents include but are not limited to titanium dioxide (in the form of anatase and/or rutile and/or brookite), zinc oxide, tin oxide, cadmium sulphide, tungsten trioxide and molybdenum trioxide. Alternatively, combinations of two or more of these agents may be used. In a preferred embodiment the agent is titanium dioxide.

30

In the present invention, the photocatalytic composition further comprises a metal complex sensitizer. The central atom of such sensitizers can be but is not limited to

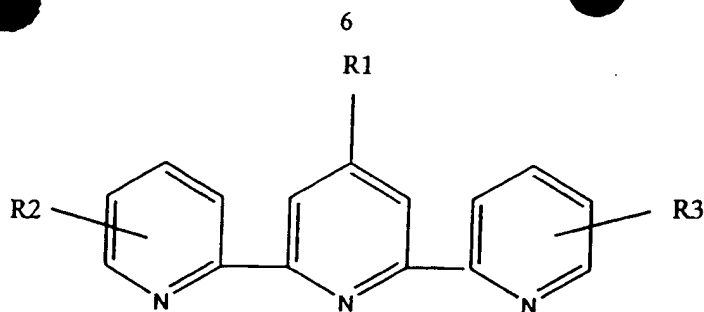
ruthenium, platinum, palladium, iridium, rhodium, osmium, rhenium, iron or copper, titanium or zinc. In one embodiment suitable sensitising agents include but are not limited to heterocyclic complexes which contain polypyridine, macrocyclic or phthalocyanine ligands and optionally other ligand types wherein at least one of the  
5 nitrogen groups is displaced by other donor groups. In a preferred embodiment of the invention the complex is any one or more of ruthenium II, III or IV or mixed oxidation state chelating complexes containing nitrogen donor atoms or a ruthenium(II), (III), (IV) or a mixed oxidation state polypyridine complex.

10 In a further embodiment the sensitising agent includes any one or more of the following groups: terpyridyls, bipyridyls, phthalocyanines, porphyrins, tetra-aza-annulenes, pyrazines, phenanthrolines and derivatives thereof and compounds with substantially similar nitrogen based ring systems.

15 These groups may be derivatised to produce compounds containing positively-charged binding sites suitable for attachment to semiconductors. Thus the sensitising agent may further include any one or more of  $R_4N^+$  or  $R_4P^+$  groups wherein each R group may be the same or different and is any one or more of the following groups: hydrogen, halogen, amine, alkyl, aryl, arylalkyl, alkoxy, heterocyclic groups, or  
20 derivatives thereof, including acid and ester derivatives, any of which may be branched or unbranched, substituted or unsubstituted,

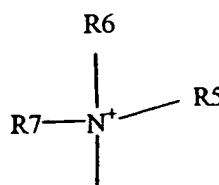
In this way sensitising agents are specifically designed wherein the molecular structure functions in combination with semiconductors where the desired operating  
25 condition is such that the un-coated semiconductor surface presents adsorption sites with a negative charge. This will occur for instance where the composition containing said agent is of alkaline pH.

On one embodiment the sensitising agent may include a terpyridal group of  
30 general formula I shown below:



(I)

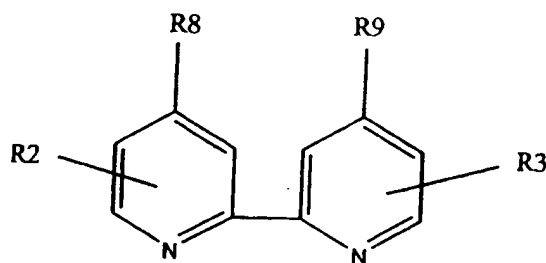
Where at least one of R1, R2 and R3 are positively charged groups which has the general formula II shown below:



(II)

Where R5-R7 are any one or more of the following groups: hydrogen, halogen, amine, alkyl, aryl, arylalkyl, alkoxy, heterocyclic groups, or derivatives thereof, including acid and ester derivatives, any of which may be branched or unbranched, substituted or unsubstituted,

Alternatively, or in addition the sensitising agent may include a bipyridyl group having the general formula III shown below:



(III)

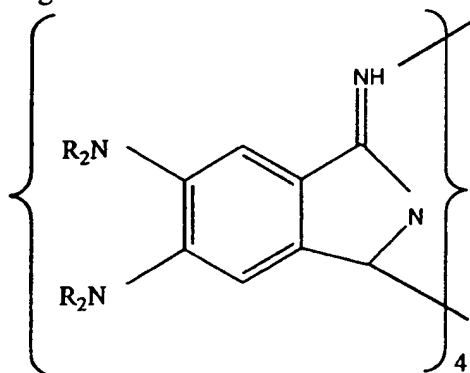
Where R8 and R9 can be the same or different and is any one or more of the following groups: hydrogen, halogen, amine, alkyl, aryl, arylalkyl, alkoxy, heterocyclic groups, or derivatives thereof, including acid and ester derivatives, any of which may be branched or unbranched, substituted or unsubstituted,

R<sub>2</sub> may be the same or different from R<sub>3</sub> and is any one or more of the following groups: hydrogen, halogen, amine, alkyl, aryl, arylalkyl, alkoxy, heterocyclic groups, or derivatives thereof, including acid and ester derivatives, any of which may be branched or unbranched, substituted or unsubstituted,

5

Alternatively, or in addition sensitising agents of the present invention may include phthalocyanines of general formula IV below:

10

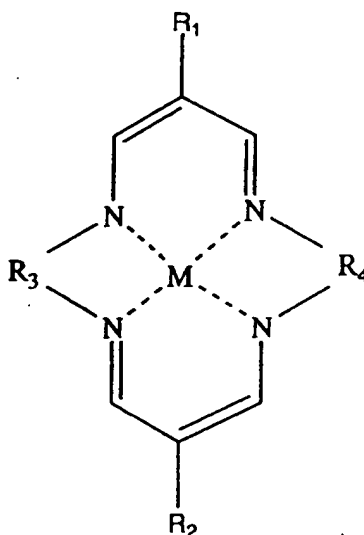


(IV)

15 Where each R group may be the same or different and is any one or more of the following groups: hydrogen, halogen, amine, alkyl, aryl, arylalkyl, alkoxy, heterocyclic groups, or derivatives thereof, including acid and ester derivatives, any of which may be branched or unbranched, substituted or unsubstituted,.

20 Alternatively, or in addition sensitising agents may include tetra-aza-annulenes (TADAs) of general formula V shown below.

25



30

(V)

R<sub>1</sub>-R<sub>4</sub> may be the same or different and is any one or more of the following groups: hydrogen, halogen, amine, alkyl, aryl, arylalkyl, alkoxy, heterocyclic groups, or



derivatives thereof, including acid and ester derivatives, any of which may be branched or unbranched, substituted or unsubstituted,.

Alternatively, or in addition the bipyridyl compounds tris(2,2'-bipyridyl - 4,4'-dicarboxylate) ruthenium (II) dichloride and tris(2,2'-bipyridyl) ruthenium (II) dichloride  
5 can be dimerised using pyrazine derivatives such as pyrazine, pyrimidine and 4, 4'-bipyridyl linking ligands using procedures well known in the art. Again as previously discussed these will be most suitable for use in operating conditions such that the uncoated semiconductor presents absorption sites with a negative charge

10

Compositions of the present invention will most preferably be in the form of a liquid. They may also be in the form of an emulsion, suspension, or in particulate form. Preferably, the light absorbing agent will comprise no more than 50% w/v of the photocatalytic composition, more preferably the light absorbing agent will comprise no  
15 more than 10% w/v of the photocatalytic composition. More preferably still the light absorbing agent will comprise no more than 1% w/v of the photocatalytic composition. Yet more preferably the light absorbing agent will comprise no more than 0.1% w/v of the photocatalytic composition. Preferably the sensitising agent will comprise no more than 1% w/v of the photocatalytic composition. More preferably the sensitising agent  
20 will comprise no more than 0.1% w/v of the photocatalytic composition.

The compositions of the present invention are effective at a whole range of pH values from 1 to 14. For compositions comprising sensitising agents of the present invention which contain polypyridine, macrocyclic or phthalocyanine ligands and  
25 optionally other ligand types wherein at least one of the nitrogen groups is displaced by other donor groups, in particular any one or more of: sensitising agent is ruthenium II, III or IV or mixed oxidation state chelating complexes containing nitrogen donor atoms, or a ruthenium(II), (III), (IV) or a mixed oxidation state polypyridine complex, then these compounds perform most effectively at pHs corresponding to a positive charged  
30 surface-state of the semiconductor component e.g for titania this corresponds to a pH of less than 7. Thus in a preferred embodiment of this aspect of the invention a composition comprising sensitising agents described above and also titania preferably has a pH of less than 7, even more preferably of less than 6, more preferably still of less than 5.

For compositions comprising a sensitising agent according to the present invention which includes any one or more of the following groups: terpyridyl, bipyridyls, phthalocyanines, porphyrins, tetra-aza-annulenes, pyrazines, phenanthralines and derivatives thereof and compounds with substantially similar nitrogen based ring systems, and may further include any one or more of  $R_4N^+$  or  $R_4P^+$  groups wherein each R group is as hereinbefore described., the preferred pH of the composition corresponds to the value where the semi-conductor component has a negatively charged surface. For titania this is pH 7 or greater. Even more preferred is a pH of greater than 8, more preferred still a pH of greater than 9.

10

It is also the case that even where the semiconductor component has a surface-excess of positive charge at a particular pH, negatively charged sites for binding positively charged sensitising agents may well be present so that both-charge types of sensitizer may effectively be used. Similarly, for systems where mixed semiconductor components e.g. titania with zinc oxide, are used both charge types of sensitising agents may be employed.

In a further aspect the present invention provides a sensitising agent which includes any one or more of the following groups : terpyridyl, bipyridyl, phthalocyanine, porphyrins, tetra-aza-annulenes, pyrazines, phenanthrolines and derivatives thereof and compounds with substantially similar nitrogen based ring systems.

The sensitising agents listed above further includes any one or more of  $R_4N^+$  or  $R_4P^+$  Where R5-R7 are any one or more of the following groups: hydrogen, halogen, amine, alkyl, aryl, arylalkyl, alkoxy, heterocyclic groups, or derivatives thereof, including acid and ester derivatives, any of which may be branched or unbranched, substituted or unsubstituted,. These groups may be derivatised to produce compounds containing positively-charged binding sites suitable for attachment to semiconductors as hereinbefore described.

30

In yet a further aspect, the present invention provides the use of a sensitising agent according to the present invention for the sensitisation of a light absorbing agent on a surface such that soils present on the surface are substantially broken down and/or the rate of accumulation of such soils on a surface is significantly diminished.

The term 'the rate of accumulation of soils is significantly diminished' in the context of the present invention means that the rate is significantly diminished as compared with a similar sample in which no sensitising agent has been applied.

5 The photocatalytically active composition, may be doped with an additional element which has the effect of reducing the energy required to promote an electron of the photochemically active material to the conductance band. Suitable doping agents may include but are not limited to platinum, palladium, cobalt, silver, copper, nickel or iron, tungsten, chromium. These may be present as the metals themselves, and/or as  
10 complexes and/or compounds thereof.

Compositions of the present invention may further include a wetting agent which may be any one or more of the following: Igepal® CA-520 [polyoxyethylene(5) isooctylphenyl ether], Igepal® CA-630 [(octylphenoxy)polyethoxyethanol], Igepal®  
15 CA-730 [polyoxyethylene(12) isooctylphenyl ether]. Preferably the concentration used will be between 0.5-5.0 wt%, even more preferably between 0.5 and 3 wt%, more preferably still between 0.5 and 2.0 wt%.

The photocatalytic compositions and/or sensitising agents of the present  
20 invention can be used in conjunction with those conventional ingredients of cleaning materials known to those skilled in the art. These may include but are not limited to water, anionic, non-ionic or amphoteric surfactants. Grease cutting, surfactant synergistic or other solvents may also be included as may antibacterial agents, suspending agents, colourants, perfumes, thickeners, preservatives and so on. Some or all of the ingredients  
25 may be of high volatility whereby a residue of photochemically active material can be left behind on a surface in a controlled manner.

The sensitising agent, or compositions according to the present invention may be applied to the surface in any appropriate form such as, for example, a liquid, cream,  
30 mousse, emulsion, microemulsion or gel form and may be dispensed either directly from the bottle or by means of for example an aerosol, pump action dispenser. These means will be known to those in the art.

One skilled in the art will appreciate that generally the compositions and/or sensitising agent according to the present invention once deposited on the surface should be substantially imperceptible to the user. This may be achieved by using materials, agents and compositions with a microscopic particle size. The microscopic particle size  
5 also aids in achieving a uniform dispersion throughout the materials and/or compositions thus maximising the efficiency of the photochemical reaction. Preferably the particle size is less than 100nm, more preferably the particle size is less than 50nm and more preferably still it is less than 20nm

10 In some circumstances one skilled in the art will appreciate the need for the photocatalytic composition and/or sensitising agent to possess larger particle sizes.

The invention will now be described with reference to the following examples in which are in no way limiting of the invention, and in which:

15

Figure 1 represents the UV/Visible spectra of the  $\lambda$  max of the target dye Gentian Violet disappearing with time as described in example 6. Horizontal axis is wavelength in nm. Vertical axis is absorbance in units measured using UV spectrometer (UV / vis spectrometer-UV 4-UNICAM)

20 ♦ represents sensitised TiO<sub>2</sub>+0.1ml Gentian Violet dye (GV) at T=0mins, ■ Sensitised TiO<sub>2</sub>+0.1ml GV T=30mins, ▲ Sensitised TiO<sub>2</sub>+0.1ml GV T=1hour, × Sensitised TiO<sub>2</sub>+0.1ml GV T=2hours, - Sensitised TiO<sub>2</sub>+0.1ml GV T=3hours, ● Sensitised TiO<sub>2</sub>+0.1ml GV T=4hours.

25 Figure 2 represents the activity of TiO<sub>2</sub> sol (sol 1) as described in example 8. Horizontal axis represents time and the vertical axis represents the change in absorbance measured using a (UV / vis spectrometer-UV 4-UNICAM)

30 ♦ represents the activity of sensitised TiO<sub>2</sub> at pH 3.28, ■ Activity of sensitised TiO<sub>2</sub> at pH 2.08, ▲ Activity of sensitised TiO<sub>2</sub> at pH 2.72, × Activity of sensitised TiO<sub>2</sub> at pH 4.02.

Figure 3 represents the activity of TiO<sub>2</sub> sol (sol 2) as described in example 8. Horizontal axis represents time and the vertical axis represents the change in absorbance measured using a (UV / vis spectrometer-UV 4-UNICAM)

- ◆ represents the activity of sensitised TiO<sub>2</sub> at pH 2.00, ■ Activity of sensitised TiO<sub>2</sub> at pH 2.64, ▲ Activity of sensitised TiO<sub>2</sub> at pH 4.12, × Activity of sensitised TiO<sub>2</sub> at pH 3.39, \* Activity of sensitised TiO<sub>2</sub> at pH 5.00, ● Activity of sensitised TiO<sub>2</sub> at pH 5.98.

Figure 4 represents the activity of TiO<sub>2</sub> sol (sol 3) as described in example 8. Horizontal axis represents time and the vertical axis represents the change in absorbance measured using a (UV / vis spectrometer-UV 4-UNICAM)

- ◆ represents the activity of sensitised TiO<sub>2</sub> at pH 4.1, ■ Activity of sensitised TiO<sub>2</sub> at pH 3.2, ▲ Activity of sensitised TiO<sub>2</sub> at pH 2.7, × Activity of sensitised TiO<sub>2</sub> at pH 2.1, \* Activity of sensitised TiO<sub>2</sub> at pH 5.2, ● Activity of sensitised TiO<sub>2</sub> at pH 6.0, - Activity of sensitised TiO<sub>2</sub> at pH 6.5-7.0.

Figure 5 represents the activity of TiO<sub>2</sub> sol (sol 4) as described in example 8. Horizontal axis represents time and the vertical axis represents the change in absorbance measured using a (UV / vis spectrometer-UV 4-UNICAM)

- ◆ represents the activity of sensitised TiO<sub>2</sub> at pH 2.74, ■ Activity of sensitised TiO<sub>2</sub> at pH 2.12, ▲ Activity of sensitised TiO<sub>2</sub> at pH 3.38, × Activity of sensitised TiO<sub>2</sub> at pH 4.00.

Figure 6 represents the activity of sensitised TiO<sub>2</sub> sol at different pH as described in example 9. Horizontal axis represents time and the vertical axis represents the change in absorbance measured using a (UV / vis spectrometer-UV 4-UNICAM)

- ◆ represents the activity of sensitised solution (1) at pH 6.7, ■ Activity of sensitised TiO<sub>2</sub> at pH 5.2, ▲ Activity of sensitised TiO<sub>2</sub> at pH 8.8.

Figure 7 represents the effect of the light source in photocatalytic activity as described in example 11. Horizontal axis represents time and the vertical axis represents the change in absorbance measured using a (UV / vis spectrometer-UV 4-UNICAM)

- ◆ represents the ratio: TiO<sub>2</sub>:Ru:1:6, ■ Ratio:TiO<sub>2</sub>: Ru:1:4, ▲ Ratio:TiO<sub>2</sub>:Ru:1:2.

### EXAMPLES

In the following examples OHP represents overhead projector, GV stands for Gentian Violet dye.

#### EXAMPLE 1

5 A nanocrystalline titanium dioxide sol was applied to the surface of a previously cleaned glass microscope slide by spin coating 0.5ml of the titanium dioxide sol at 1500rpm for 30 seconds. The glass slide was then fired at 450°C for 30 minutes. Once  
10 cool the process was repeated two further times to give 3 coats of the nanocrystalline titanium dioxide. The slide was then immersed in an aqueous  $1 \times 10^{-6}$ M solution of tris(2,2'-bipyridyl-4,4'-dicarboxylate)Ru(II)(dichloride) for 30 minutes to allow adsorption of the sensitising agent to the titanium dioxide. The slide was removed, rinsed with water to remove any unbound ruthenium complex and then stained with a 0.3%  
15 Gentian Violet solution (N-4[Bis[4-dimethylamino)-phenyl]methylene]-2,5-cyclohexadien-1-ylidene]-N-methylmethanaminium chloride) in 20% ethanol by immersing in the dye for 5 minutes. Once again the slide was washed with water to remove any unbound dye. The whole process was repeated once more for a second identical slide and then twice more, but omitting immersing these two slides in the  
20 tris(2,2'-bipyridyl-4,4'-dicarboxylate)Ru(II)(dichloride) to give two un-sensitized control slides.

One each of the sensitised and un-sensitised slides was kept in total darkness. The other two slides were placed on top of an overhead projector fitted with two 24V  
25 250W tungsten halogen bulbs. Decolourisation of the purple colour was monitored as the Gentian Violet was decomposed photocatalytically. Within 50 minutes there was a noticeable difference between the colour of the sensitised and un-sensitised slides particularly when compared to the controls stored in the dark. On the slide with the sensitised titanium dioxide the Gentian Violet was decomposing and hence the purple  
30 colour fading. This continued until there was no purple colour left. Decomposition of the Gentian Violet did also occur on the un-sensitised slide but at a significantly slower rate.

## EXAMPLE 2

A nanocrystalline titanium dioxide sol thickened with methylcellulose was screen printed on to a series of cleaned glass microscope slides. The printed titanium dioxide  
5 films were then fired at 450°C for 30 minutes. Half of the slides were then immersed in an aqueous  $1 \times 10^{-6}$  M solution of tris(2,2'-bipyridyl-4,4'-dicarboxylate)Ru(II)(dichloride) for 30 minutes to allow adsorption of the sensitising agent to the titanium dioxide. The slides were then removed from the sensitising solution and washed with water to remove any unbound ruthenium complex. All the slides, both sensitised and unsensitised  
10 were then divided into two groups. One set was immersed in 0.3% Gentian Violet solution (N-4-[Bis[4-dimethylamino)-phenyl]methylene]-2,5-cyclohexadien-1-ylidene]-N-methylmethanaminium chloride) in 20% ethanol for 5 minutes and the second set into a 0.3% aqueous solution of Acid Orange dye (4-[(2-hydroxy-1-naphthalenyl)azo]-benzenesulfonic acid monosodium salt) for 5 minutes. A sensitised and unsensitised slide  
15 dyed with either the Gentian Violet or Acid Orange stains was placed in total darkness and used as a control for each treatment. A second equivalent set was left exposed to daylight next to the window of a south-facing window. A third and final set was also left exposed to the daylight through a south-facing window but these slides were covered with a 6mm thick piece of Perspex which substantially absorbs the UV component of the  
20 light. Decolourisation of both the purple and orange colours was monitored as both the Gentian Violet and Acid Orange were decomposed photocatalytically. After 48 hours exposure to light the slides dyed with Gentian Violet and left directly on the open bench were partially decolourised. The slides stored under the Perspex had begun to decolourize but at a slower rate than those not under Perspex. By day 7 the dye on all the  
25 slides left just on the bench had either completely or almost completely disappeared. The slides under Perspex reached the same amount of decolourization on day 14. There was no change in the colour of the slides stored in the dark. All the light exposed slides were re-dyed with either Gentian Violet or Acid Orange and treated exactly as before. None of the Acid Orange stained slides would re-stain. The sensitised titanium dioxide slides  
30 stained with Gentian Violet and exposed directly to daylight decolourised completely within 24 hours. The unsensitised slide had still not completely decolourised 5 days after restaining. The slides under Perspex were still coloured 5 days after restaining, however the sensitised slide had faded to a greater extent than the unsensitised slide.

### EXAMPLE 3

To 1.0ml of a nanocrystalline titanium dioxide solution, 4.0ml of a  $3.4 \times 10^{-5}$  M aqueous solution of tris(2,2'-bipyridyl-4,4'-dicarboxylate)Ru(II)(dichloride) were added and the resulting solution mixed well using a vortex mixer. A film of this solution was prepared by spin coating 0.1ml of this solution at 1500rpm on a clean glass microscope slide for 30 seconds. The film was dried using a hand held hot air drier and the process repeated twice more to give a total of 3 coats on the microscope slide. A second slide was then prepared in exactly the same way. Both slides were then immersed into a solution of 0.3% Gentian Violet (N-4-[Bis[4-dimethylamino)-phenyl]methylene]-2,5-cyclohexadien-1-ylidene]-N-methylmethanaminium chloride) in 20% ethanol for 5 minutes. The slides were removed, rinsed with water to remove any excess stain and allowed to air dry. One slide was kept in total darkness and the second was placed on top of a piece of 6mm thick Perspex (to remove any UV light) on an overhead projector fitted with two 24V 250W tungsten halogen bulbs. The purple colour on the light exposed slide steadily decomposed and after 3 hours had completely faded. There was no change in the colour of the slide stored in darkness.

### EXAMPLE 4:

#### Titania sols preparation

Kormann method. (C. Kormann, D.W. Bahnemann, M.R. Hoffmann, J. Phys. Chem., 1988, 92, 5196)

TiCl<sub>4</sub> (3.5ml) was slowly added to cold de-ionised water (900ml) under vigorous stirring. The resulting clear solution was stirred at 0°C for 3 hours then dialysed between 2 hours and 24 hours. The clear solution is then dried using a rotary-evaporator (Temperature of water bath=30°C). The resulting white powder (TiO<sub>2</sub>) is then re-suspended into de-ionised water at the desired concentration. The dialysis membrane-Visking-Size1350/1 MWCO 1350 Daltons was treated prior to use. The membrane was left for 30 minutes at 80°C in a solution containing EDTA (1mM) and 2% NaHCO<sub>3</sub>. The membrane was then washed thoroughly with de-ionised water.



Method according to GB 1 412 937

An aqueous  $\text{TiCl}_4$  solution (50ml of  $\text{TiCl}_4$  diluted in 500ml de-ionised water) was added into a beaker containing de-ionised water (3L) and concentrated ammonia (40ml) with continuous stirring. The white mixture was stirred for about 20 minutes then allowed to settle. The supernatant was removed using a peristaltic pump. The volume was completed again to 3L with de-ionised water, stirred then allowed to settle. The supernatant was removed. This process was repeated twice. The volume was completed with de-ionised water to 3.5L. The mixture was stirred, the pH was checked (pH 8.8) then a nitric acid solution (1M) was added slowly to get pH close to 3.3. The mixture was stirred for 30-45 minutes then allowed to settle. The supernatant was removed and the volume was made up with de-ionised water to 3-3.5L. The mixture was washed until the conductivity was below  $500\mu\text{S}$ . The supernatant was removed then nitric acid (1M, 23.2ml) was added to the white mixture. The mixture was stirred for about 20 minutes then was left to age for about a week. In order to increase the peptisation step, the mixture can be heated gently to  $60-70^\circ\text{C}$  for 30 minutes then allowed to settle.

Isopropoxide route.

Titanium-isopropoxide (Aldrich, 400ml, 97%) was added rapidly to a beaker containing de-ionised water (1L). The precipitated  $\text{TiO}_2$  was decanted and washed 4 times with de-ionised water (4x500ml) then filtered. The wet filtered solid was digested at  $70^\circ\text{C}$  with concentrated nitric acid (16.7ml) and de-ionised water (volume total 800ml) for 30 to 1h 30min to produce a sol.

**EXAMPLE 5**Synthesis of Tris(2,2'-dipyridyl-4,4'-dicarboxylate)ruthenium(II)(dichloride)

$\text{RuCl}_3 \cdot x\text{H}_2\text{O}$  (1.33mmol Ru), 1-methyl-2-pyrrolidinone (15ml) and 2,2'-dipyridyl-4,4'-dicarboxylate (4.1mmol) were added into a round bottomed flask and then purged with Ar or  $\text{N}_2$ . The mixture was heated to reflux in the dark for 1h30min. 1-methyl-2-pyrrolidinone (25ml) was added to the flask and the reflux was continued for a

further 2 hours under Ar or N<sub>2</sub>. The mixture was allowed to cool to room temperature and kept under Ar or N<sub>2</sub> overnight. The dark mixture was filtered. The resulting reddish brown solid was washed with 1-methyl-2-pyrrolidinone (2x20ml) and diethyl ether (3x20ml) then dried under vacuum. Yield = 0.61g. Product contained 1.2moles of 1-methyl-2-pyrrolidinone.

### EXAMPLE 6

#### Activity test

A mixture of TiO<sub>2</sub> sol (1ml, 1g/L) and tris(2,2'-dipyridyl-4,4'-dicarboxylate) ruthenium(II)(dichloride) (4ml,  $c=3.4 \times 10^{-5}M$ ) was stirred for about 1 minute using a rotamixer. The pH was recorded then the target dye gentian violet (GV dissolved in 20% ethanol solution, 0.05ml or 0.08ml, 0.03 wt/v%) was added and the mixture was stirred again. The initial colour of the sensitised TiO<sub>2</sub> sol was yellow. Addition of gentian violet produced a purple colour at pH 3 and higher. A UV/Visible spectrum was taken at this stage. The vial containing the mixture was placed onto an overhead projector (2cm height from the glass, in order to reduce heat). A UV/Visible spectrum was used to observe the colour change over a period of time at lamda max of the Target dye (Gentian violet or crystal violet) = 588nm in the white light spectrum. (OHP used: Model Ensign. Lamp:24V-250W-3860 lux)

In order to get quantitative data a UV/Visible spectrum was taken at different times. The results are shown in figure 1.

### EXAMPLE 7

#### Effect of different steps of a preparation, effect of peptisation and pH Effect.

The "activity" has been tested at different stages of the preparation of the TiO<sub>2</sub> by hydrolysis of TiCl<sub>4</sub> (Kormann Method-see example 4 for details). The effect of peptisation has also been looked at. From the results shown in Table 1, the peptisation as well as the particle size seemed to have little effect on the activity.

The details of the different preparations as well as the activity test are summarised in Example 4 and 6.

Code	Ti Source	TiO <sub>2</sub> sol Preparation	Temp of Peptisation	Time for Peptisation	Particle size (nm)	Time to decolourise 0.05ml GV
Sol A	TiCl <sub>4</sub>	Prepared as per example 1 of GB 1412 937 (Woodhead)	60-65C for 30mins	99% after 48hours	22.5	30min
Sol B	Ti isopropoxide	Isopropoxide route	Ambient	99% after 7 days	720.7	90min
Sol C	Ti isopropoxide	Isopropoxide route	60C for 30mins	* 95% after 7 days	34.5	40min
Sol D	Ti isopropoxide	Isopropoxide route	70C for 30mins		95.4	55min
Sol E	TiCl <sub>4</sub>	Prepared as per example 1 of GB 1412 937 (Woodhead)	Ambient	99% after 7 days	720.3	60min

- 5 \*Significant amount of non-dispersed material present after 5 days so further nitric acid (1M) was added to give NO<sub>3</sub><sup>-</sup>:Ti ca 0.27

Note: the mixture was 1ml TiO<sub>2</sub> (1g/L) and 3ml sensitiser (c=3.4x10<sup>-5</sup>M)

Table 1. Effect of peptisation on activity.

**EXAMPLE 7b**

The pH of the sensitised sol (TiO<sub>2</sub> sol prepared by Kormann method) was found to be different at each step of the process. The results are summarised in Table 2. The pH  
5 was measured using a pH meter (HANNA Instruments- HI8424 microcomputer).

Steps of the process	pH of the sensitised sol	Time to decolourise 0.08ml Gentian Violet
Before dialysis	1.85	2h45min
After dialysis	2.95	2h20min
Sol dried using rotary-evaporator then re-suspended	3.16	4h40min
Sol dried using freeze-drier then re- suspended *	3.08	2h45min mixture still purple.

\* TiO<sub>2</sub> was very difficult to re-suspend. After sonication the sol was cloudy.

10

Table 2. Effect of different steps of the process on activity.

The results indicate that the activity may be related to pH.

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**EXAMPLE 8**

The effect of pH on activity of various sols

Four different sols have been tested at pH ranging from 2 to 7. They are:

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-(Sol 1) Hydrolysis of  $\text{TiCl}_4$  followed by a dialysis, dried on rotary-evaporator then re-suspended. The pH of the sol was adjusted with HCl (1M) or NaOH (0.01M).

-(Sol 2) Hydrolysis of  $\text{TiCl}_4$  followed by a dialysis only. The pH of the sol was adjusted with HCl (1M) or NaOH (0.01M).

5   -(Sol 3) Precipitation of titanium –isopropoxide followed by peptisation with nitric acid. The pH of the sol was adjusted with  $\text{HNO}_3$  (0.1M) or NaOH (0.01M).

-(Sol 4) Precipitation of  $\text{TiCl}_4$  followed by preptisation with nitric acid. The pH of the sol was adjusted with  $\text{HNO}_3$  (0.1M) or NaOH (0.01M).

10       This experiment was carried out over a 2 hour period. All sols were prepared at 1g/L and the amount of Gentian Violet (0.08ml, 0.03 wt/v%) was kept the same for each type of sol. Most of the sensitised sols had a precipitate or were precipitating. However, the systems were still working even in presence of a precipitate. The “activity” was reduced with an increase of pH. The results are summarised in Table 3.

15       A mixture of  $\text{TiO}_2$  sol (1ml, 1g/L) and tris(2,2'-dipyridyl-4,4'-dicarboxylate) ruthenium(II)(dichloride) (4ml,  $c=3.4 \times 10^{-5}\text{M}$ ) was stirred for about 1 minute using a rotamixer. The pH was adjusted then the target dye gentian violet (GV dissolved in 20% ethanol solution, 0.08ml, 0.03 wt/v%) was added and the mixture was stirred again. The  
20   initial colour of the sensitised  $\text{TiO}_2$  sol was yellow. Addition of gentian violet produced a purple colour at pH 2.5 and higher. Below pH 2.5, addition of gentian violet produced a blue-green colour. A UV/Visible spectrum was taken at this stage. The vial containing the mixture was placed onto an overhead projector (2cm height from the glass, in order to reduce heat). A UV/Visible spectrum was used to observe the colour change over a  
25   period of time. (OHP used: Model Ensign. Lamp:24V-250W-3860 lux). In order to get quantitative data a UV/Visible spectrum was taken at different times.

Sol preparation	pH of sensitised sols	Time to decolourise 0.08ml Gentian Violet	Activity rank
1	2.08	The mixture was nearly yellow and clear after 2 hours.	High

	2.72	The mixture was orange-pink with a slight precipitate after 2 hours.	Medium
	3.28	The mixture was still purple with a precipitate after 2 hours.	Low
	4.02	The mixture was orange-pink with a precipitate after 2 hours.	Medium
2	2.00	The mixture was green-yellow with no real precipitate after 2 hours.	High
	2.64	The mixture was orange-pink with a precipitate after 2 hours.	Medium
	3.39	The mixture was orange-pink with a precipitate after 2 hours.	Medium
	4.12	The mixture was orange-pink with a medium precipitate after 2 hours.	Medium
	5.00	The mixture was purple with a Low precipitate after 2 hours.	Low
	5.98	The mixture was purple with a Low precipitate after 2 hours.	Low
3	2.1	The mixture was yellow with a High precipitate after 25 minutes.	High
	2.7	The mixture was yellow with a High precipitate after 35 minutes.	High
	3.2	The mixture was yellow with a High precipitate after 1h30min	High
	4.1	The mixture was yellow with a High precipitate after 2 hours.	High
	5.2	The mixture was yellow with a High precipitate after 2 hours.	High
	6.0	The mixture was orange-pink with a Medium precipitate after 2 hours.	Medium
	7.6	The mixture was purple with a Low precipitate after 2 hours.	Low

4	2.12	The mixture was yellow after 30 minutes.	High
	2.74	The mixture was yellow with a precipitate after 25 minutes.	High
	3.38	The mixture was yellow with a precipitate after 1h30min.	High
	4.00	The mixture was yellow with a precipitate after 2 hours.	High

Table 3. pH effect on the activity of different type of sols.

The results are also shown in Figure, Figure 3, and Figure 4, and Figure 5.

5

### EXAMPLE 9

#### Addition of stabilisers

Buffer solutions were obtained by diluting the powder buffer (BDH chemicals) into the required amount of de-ionised water.

Attempts to stabilise  $\text{TiO}_2$  sols prepared from the hydrolysis of  $\text{TiCl}_4$  (Kormann method) were made using buffer solution pH 7 and pH 9.2. Addition of buffer solution pH 7 into a  $\text{TiO}_2$  sol (10ml, 1g/L) produced a precipitate at pH 7. Addition of a buffer solution pH 7 or pH 9.2, de-ionised water and solid  $\text{TiO}_2$  produced cloudy solutions at different pHs (see Table 4).

15

No solutions	Sols Compositions	pH	Particle size (nm)	Observations
1	2ml buffer pH7 10ml water 9-10mg $\text{TiO}_2$	6.8	208nm	Cloudy. No apparent precipitate.

2	2ml buffer pH7 8ml water 9-10mg TiO <sub>2</sub>	6.9	120nm	Cloudy. No apparent precipitate.
3	1ml buffer pH7 10ml water 9-10mg TiO <sub>2</sub>	6.5	121nm	Cloudy. No apparent precipitate.
4	1ml buffer pH7 9ml water 9-10mg TiO <sub>2</sub>	5.2	189nm	Cloudy. No apparent precipitate.
5	0.5ml buffer pH7 9.5ml water 9-10mg TiO <sub>2</sub>	5.2		Cloudy. A precipitate was observed after 1h30min.
6	1ml buffer pH9.2 10ml water 9-10mg TiO <sub>2</sub>	8.2		Cloudy.
7	2ml buffer pH9.2 10ml water 9-10mg TiO <sub>2</sub>	8.8		Cloudy.

Table 4. Buffer solutions addition effect.

- 5 Solutions (1) and (4) were found to be cloudier than solutions (2) and (3). The particle size was higher for (1) and (4) this may correspond to the cloudiness of the solutions. After 24 hours, a slight precipitate was observed in solutions (1) and (4).

Attempts were made to increase pH with sodium hydroxide solutions failed.

- 10 Addition of acetylacetonate to a TiO<sub>2</sub> sol produced a stable sol with a pH of 2.3-2.4. Increasing the pH with a sodium hydroxide solution produced precipitation at pH around 7.



The activity was tested for the sensitised solution (1), sensitised  $\text{TiO}_2$  sols at pH 5.0 and 8.8 (note: the pH was increased by addition of NaOH (0.01M)). The target dye decolourised quicker at pH 5 although for all three samples there was some target dye left not decolourised after 2 hours. The results are summarised in figure 6.

5

Poly(vinyl alcohol) (PVA) was tested as a potential stabiliser for  $\text{TiO}_2$  sols. It was found that addition of a large excess or too little caused precipitation of the sols when the pH was increased with sodium hydroxide. PVA can be dissolved by sonication or by gentle heating in water then can be added to a  $\text{TiO}_2$  sol. Addition of PVA directly to a  $\text{TiO}_2$  sol, produced a precipitate.

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The activity of a sensitised sol containing PVA at pH around 3 and 7 was tested. The results show that PVA and an increase in pH slowed down the activity.

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#### EXAMPLE 10

##### $\text{TiO}_2$ :Sensitiser ratio effect

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The ratio  $\text{TiO}_2$ :Sensitiser or  $\text{TiO}_2$ :Ru has been looked at for a particular  $\text{TiO}_2$  sol (Kormann method,  $\text{TiO}_2$  sol dialysed only). The sol tested was obtained from hydrolysis of  $\text{TiCl}_4$  followed by a dialysis. The experiment involved variation of ruthenium and kept the  $\text{TiO}_2$  fixed. The target dye was decolourised quicker in 1:6  $\text{TiO}_2$ :Ru ratio than in 1:2  $\text{TiO}_2$ :Ru ratio

25

The target dye gentian violet (0.05ml, 0.03 wt/v%) was decolourised within 3 hours in 1:6  $\text{TiO}_2$ :Ru ratio whereas in 1:4  $\text{TiO}_2$ :Ru ratio and 1:2  $\text{TiO}_2$ :Ru ratio the gentian violet decolourised within 4 and 5 hours, respectively.

#### EXAMPLE 11

Light source Effect.

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A mixture of  $\text{TiO}_2$  sol (1ml, 1g/L) and tris(2,2'-dipyridyl-4,4'-dicarboxylate) ruthenium(II)(dichloride) (3ml,  $c=3.4 \times 10^{-5}\text{M}$ ) was stirred for about 1 minute using a

rotamixer. Addition of gentian violet (0.05ml, 0.03 wt/v%) produced a purple colour. A UV/Visible spectrum was taken at this stage. The vial containing the mixture was placed onto an overhead projector (2cm height from the glass, in order to reduce heat) or in a light box. UV/Visible spectra were used to observe the colour change over a period of time. Normally all the activity work was done using an overhead projector as a light source. Other light sources such as daylight bulbs (40W and 100W), tungsten filament tube (35W), tungsten filament bulb (100W) and fluorescent tube (8W) have also been investigated during this study. The results showed that the process still works with all light sources tried albeit much slower using daylight or tungsten filament bulbs than the light from an overhead projector because of the higher intensity of the overhead projector.

The results are shown in Figure 7

#### EXAMPLE 12:

A range of dyes have been tested as potential sensitising agent. They include: copper or iron complexes containing sulfonated phthalocyanine ligands, silicon complex containing phthalocyanine ligand and ruthenium complexes containing bipyridyl or functionalised bipyridyl complexes (e.g: carboxylate, phosphonate) ligands and anions (e.g.: Cl, NCS) and rose bengal.

A mixture of TiO<sub>2</sub> sol (1ml, 1g/L) and sensitizing agent (3ml,  $c=3.4 \times 10^{-5}M$ ) was stirred for about 1 minute using a rotamixer. Addition of gentian violet (0.05ml, 0.03wt/v%) was added as a target dye. A UV/Visible spectrum was taken at this stage. The vial containing the mixture was placed onto an overhead projector (2cm height from the glass, in order to reduce heat). UV/Visible spectra were used to observe the colour change over a period of time. All dyes tested decolourised the target dye gentian violet at different rates.

#### EXAMPLE 13

Activity of different TiO<sub>2</sub> sols.

Several TiO<sub>2</sub> sols have been tested for their activity including commercially available types from the Millennium Performance Chemicals, 85 Avenue Victor Hugo, 92563 Rueil-Malmaison Cedex, France. A mixture of TiO<sub>2</sub> sol (1ml, 1g/L) and tris(2,2'-dipyridyl-4,4'-dicarboxylate)ruthenium(II)(dichloride) (4ml,  $c=3.4 \times 10^{-5}M$ ) was stirred for about 1 minute using a rotamixer. Addition of gentian violet (0.08ml, 0.03wt/v%) produced a purple colour. A UV/Visible spectrum was taken at this stage. The vial containing the mixture was placed onto an overhead projector (2cm height from the glass, in order to reduce heat). UV/Visible spectra were used to observe the colour change over a period of time. The results are summarised in Table 5.

10

TiO <sub>2</sub> sol source	Time required to decolourise 0.08ml of gentian violet (0.03wt/v%)
Kormann method	>4hours
J. Woodhead Patent	1h30mins
Isopropoxide route	1h30mins
Millennium sol in acidic medium	15mins

**Table 5:** Activity of different sols.

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#### **EXAMPLE 14:**

Activity of different sols in basic medium.

The TiO<sub>2</sub> sol (made from isopropoxide route) containing PVA (MW:15,000) was prepared as follows. PVA (0.10g, MW:15,000) was diluted in hot de-ionised water (50ml) then allowed to cool to room temperature. A known amount of concentrated TiO<sub>2</sub> sol was added to the PVA solution under vigorous stirring. The volume was completed to 100ml with de-ionised water. Final TiO<sub>2</sub> concentration 1g/L.

#### **25    *Basic system no PVA.***

A mixture of TiO<sub>2</sub> sol (isopropoxide route, 10ml, 1g/L) and tris(2,2'-dipyridyl-4,4'-dicarboxylate)ruthenium(II)(dichloride) (40ml,  $3.4 \times 10^{-5}M$ ) was stirred using a

stirrer hotplate. The pH was adjusted by addition of a sodium hydroxide solution (0.1M) to pH 10. Gentian violet (0.08ml, 0.03wt/v%) was added to the mixture (volume used: 5ml).

5 ***Basic system with PVA (1).***

A mixture of TiO<sub>2</sub> sol containing PVA at pH 10.03 (1ml, 1g/L) and tris(2,2'-dipyridyl-4,4'-dicarboxylate)ruthenium(II)(dichloride) also at pH 10.1 (4ml,  $c=3.4 \times 10^{-5}$ M) was stirred for about 1 minute using a rotamixer. The pH (9.85) was adjusted with a sodium hydroxide solution (0.1M) in order to get pH 10. Gentian violet (0.08ml, 0.03 wt/v%) was added to the mixture.

***Basic system with PVA (2).***

A mixture of TiO<sub>2</sub> sol containing PVA (10ml) and tris(2,2'-dipyridyl-4,4'-dicarboxylate)ruthenium(II)(dichloride) (40ml,  $c=3.4 \times 10^{-5}$ M) was stirred using a stirrer hotplate. The pH was adjusted with a sodium hydroxide solution (0.1M) to pH 10. Gentian violet (0.08ml, 0.03 wt/v%) was added to the mixture (volume used: 5ml).

***Millennium basic system.***

A mixture of Millennium TiO<sub>2</sub> sol in Basic medium(1ml, 1g/L) and tris(2,2'-dipyridyl-4,4'-dicarboxylate)ruthenium(II)dichloride (4ml,  $c=3.4 \times 10^{-5}$ M) was stirred for about 1 minute using a rotamixer. The pH was adjusted with a sodium hydroxide solution (0.1M) to pH 10.

Addition of gentian violet (0.05ml, 0.03 wt/v%) produced a purple colour in all systems. A UV/Visible spectrum was taken at this stage. The vial containing the mixtures were placed onto an overhead projector (2cm height from the glass, in order to reduce heat). UV/Visible spectra were used to observe the colour change over a period of time.

The results are summarised in table 6.

TiO <sub>2</sub> sol source	Time required to decolourise	Remarks
	0.08ml of gentian violet	
Isopropoxide route no PVA	> 4 hours	Only stable for a short period of time.
Isopropoxide route with PVA (1)	> 4 hours	PVA helped stabilised the system.
Isopropoxide route with PVA (2)	> 4 hours	PVA helped stabilised the system.
Millennium in basic medium	1h10min	Stable sol. No need to add any surfactant.

Table 6: Activity of TiO<sub>2</sub> sols in basic medium**EXAMPLE 15:**

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A solution containing a TiO<sub>2</sub> sol (Millennium TiO<sub>2</sub> sol in basic medium, 10g/L, 5.0ml), tris(2,2'-bipyridyl-4,4'-dicarboxylate)ruthenium(II)(dichloride) (1.8ml,  $3.4 \times 10^{-5}$ M), Igepal® CO-720 (0.18g) and de-ionised water (3.2ml) was stirred for few minutes using a rotamixer. The pH was adjusted to 10 by addition of a sodium hydroxide solution (0.1M). The mixture was wrapped into some aluminium foil and left standing overnight to equilibrate. A solution containing a TiO<sub>2</sub> sol (Millennium TiO<sub>2</sub> sol in basic medium, 10g/L, 5.0ml), Igepal® CO-720 (0.18g) and de-ionised water (5.0ml) was stirred for few minutes using a rotamixer. The pH was adjusted to 10 by addition of a sodium hydroxide solution (0.1M).

15

Thin films of these solutions were prepared by spin coating 0.1ml of these solutions at 100 to 500rpm on a clean glass microscope slide for 80 seconds. The film was dried using a hot air gun and the process repeated to give a total of 2 coats on the microscope slide. A second slide was then prepared in exactly the same way. All slides were then immersed into a solution of 0.3% Gentian Violet in 20% ethanol for 5 minutes. The slides were removed, rinsed with water to remove any excess stain and allowed to air dry. One slide was kept in total darkness and the second was placed onto an overhead projector (Model Ensign. Lamp: 24V-250W-3860 lux). The purple colour on the films

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faded after 3 hours 30 min. There was no change in the colour of the slide stored in darkness.

### **EXAMPLE 16:**

- 5 The titania sols have been characterised by TEM (transmission electron microscopy). The samples were prepared by pipetting a few drops of the sol onto holey carbon films. Gold grids were used to avoid support corrosion. The microscope used was a Philips CM20, operated at 200kV. The results are summarised in Table 7.

TiO <sub>2</sub> sol source	TEM results
Prepared as per example 1 of GB 1412 937 (Woodhead) (peptised at 60°C) (A)	Irregularly shaped titania crystallites mainly anatase with a size of 10 nm.
Prepared as per example 1 of GB 1412 937 (Woodhead) (peptised at room temperature) (B)	Irregularly shaped titania crystallites mainly anatase with a size of 10 nm.
Sensitised TiO <sub>2</sub> sol* (C)	Irregularly shaped titania crystallites mainly anatase with a size of 10 nm.
Isopropoxide route (D)	Sample had to be diluted. The titania crystallites were more tightly packed together. The crystallite size is very small. The particles formed self-supporting films over holes on the carbon film.
Millennium sol Basic medium (E)	Titania particles appeared to be much more agglomerated than samples A and C. As a result of this agglomeration, the particles formed self-supporting films over holes on the carbon film.

10

\* TiO<sub>2</sub> sol (Prepared as per example 1 of GB 1412 937 (Woodhead) 10g/L, 5.0ml), tris(2,2'-bipyridyl-4,4'-dicarboxylate) ruthenium(II)(dichloride) (1.8ml, 3.4x10<sup>-5</sup>M) and de-ionised water (3.2ml)

**Table 7:** TEM results of different source of titania.

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**EXAMPLE 17:****Activity of stabilised sensitised titania versus titania only .*****TiO<sub>2</sub> only with PVA.***

- 5           The TiO<sub>2</sub> sol (made from isopropoxide route) containing PVA (MW:15,000) was prepared as follows. PVA (0.10g, MW:15,000) was diluted in hot de-ionised water (50ml) then allowed to cool to room temperature. A known amount of concentrated TiO<sub>2</sub> sol was added to the PVA solution under vigorous stirring. The volume was completed to 100ml with de-ionised water. Final TiO<sub>2</sub> concentration 1g/L.
- 10       A mixture of TiO<sub>2</sub> sol containing PVA (1ml, 1g/L) and de-ionised water (4ml) was stirred for about 1 minute using a rotamixer. Gentian violet (0.08ml, 0.03wt/v%) was added to the mixture.

***Sensitised TiO<sub>2</sub> with PVA-Acidic medium.***

- 15           The TiO<sub>2</sub> sol (made from isopropoxide route) containing PVA (MW15,000) was prepared as follows. PVA (0.10g, MW 15,000) was diluted in hot de-ionised water (50ml) then allowed to cool to room temperature. A known amount of concentrated TiO<sub>2</sub> sol was added to the PVA solution under vigorous stirring. The volume was completed to 100ml with de-ionised water. Final TiO<sub>2</sub> concentration 1g/L. A mixture of TiO<sub>2</sub> sol
- 20       containing PVA (1ml, 1g/L) and tris(2,2'-dipyridyl-4,4'-dicarboxylate)ruthenium(II)(dichloride) (4ml,  $c=4.3 \times 10^{-5}M$ ) was stirred for about 1 minute using a rotamixer. Gentian violet (0.08ml, 0.03wt/v%) was added to the mixture.

***Sensitised TiO<sub>2</sub> with PVA -Basic medium.***

- 25           A mixture of TiO<sub>2</sub> sol containing PVA (10ml) and tris(2,2'-dipyridyl-4,4'-dicarboxylate)ruthenium(II)(dichloride) (40ml,  $c=3.4 \times 10^{-5}M$ ) was stirred using a stirrer hotplate. The pH was adjusted with a sodium hydroxide solution (0.1M) to pH 10. Gentian violet (0.08ml, 0.03 wt/v%) was added to the mixture (volume used: 5ml).

- 30           A UV/Visible spectrum is taken at this stage. The vials containing the mixtures were placed onto an overhead projector (2cm height from the glass, in order to reduce heat). UV/Visible spectra were used to observe the colour change over a period of time.

TiO<sub>2</sub> without dye has significantly slower photocatalytic activity when compared with sensitised TiO<sub>2</sub>

#### EXAMPLE 18:

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##### Breakdown of 4-chlorophenol (halogenated pollutant) using TiO<sub>2</sub>.

A microscope slide containing a thin film of sensitised TiO<sub>2</sub> was added into a solution of 4-chlorophenol (99+%, Aldrich, 8ml, 10<sup>-4</sup>M). The vial containing the solution and the slide was placed onto an overhead projector. The degradation of 4-chlorophenol was monitored using UV/Visible analysis. A spectrum was taken over a period of time at max of the 4-chlorophenol (=280nm). The absorbance at 280nm was decreasing over time.

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#### EXAMPLE 19:

##### Demonstration of the photocatalytic nature of TiO<sub>2</sub> in addition to sensitiser

A mixture of TiO<sub>2</sub> sol (5ml, 10g/L), tris(2,2'-dipyridyl-4,4'-dicarboxylate)ruthenium(II)(dichloride) (1.8ml, 3.4x10<sup>-5</sup>M) and de-ionised water (3.2ml) was stirred for about 1 minute using a rotamixer. Half of the volume of the mixture was used for testing. A target dye, gentian violet (0.08ml, 0.03wt/v%) was added to the mixture producing a blue-purple colour. A UV/Visible spectrum was taken. The vial containing the mixture was placed onto an overhead projector (2cm height from the glass in order to reduce heat). A UV/Visible spectrum was taken over a period of time in order to observe the colour change. Once the spectrum showed no trace of target dye, the same amount of gentian violet was added to the same mixture. The all process was repeated twice. The target dye gentian violet was still decolourising on the third addition but at a slower rate than on the first addition.

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#### EXAMPLE 20:

Zinc oxide was prepared according to the method outlined by Bahnemann et al, J. Phys. Chem., (1987), 91, 3789. The oxide suspension (made by stirring the ZnO solid into a sodium hydroxide solution at pH9) was then sensitised with 4,4'-dicarboxate, tris



(2,2'-bipyridyl) Ru (II) dichloride according to the method outlined in previous examples. Gentian violet dye was added to both the sensitised sample and the non-sensitised control sample, and the UV/visible spectrum was recorded as a function of time under illumination with white light (5,000 lux). The results demonstrate that the absorption peak associated with Gentian Violet decreases faster with the sensitised ZnO compared to the control.

### EXAMPLE 21:

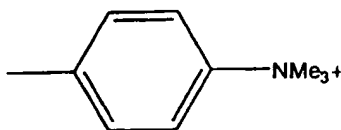
The methods described below offer general synthetic pathways to sensitising agents specifically designed to function in combination with semiconductors where the desired operating condition is such that the un-coated semiconductor surface presents a significant number of adsorption sites with a negative charge. Typical positively charged groups for use as binding sites include, but are not limited to  $R_4N^+$  groups and  $R_4P^+$  groups, where R is as hereinbefore described

#### Novel dyes based on terpyridyl

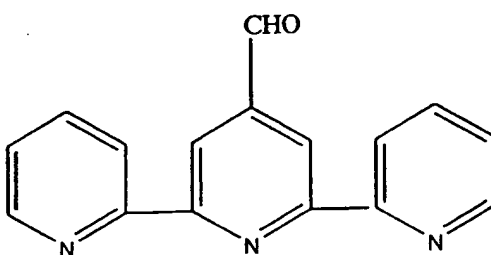
Terpyridyl-based sensitisers with phosphonate chelating ligands (Graetzel et al, WO 95/29924) and with other ligand types (Graetzel et al, WO 94/0449) have been used in conjunction with titania in dye-sensitised solar cells. The terpyridyl group of general formula I can be synthesised with e.g. R1 as a positively charged unit. One example where R5-7 of formula II are methyl, is synthesised according to procedures where the intermediate is made by the method outlined in Recl.Trav. Chim. Pays. Bas, 1959, v78, 408. This nitrated aryl group is then changed into the terpyridyl unit by the method outlined by McWhinne et al (J Organoetallic chem., 1968, v11, 499). The nitro group is then reduced to the amine by hydrazine hydrate under Pd/C catalysis followed by reaction with excess methyl iodide to form the quaternary nitrogen terpyridyl ligand desired.

Yet another variant of the positively charged terpyridyl molecule [described by general formula I] can be synthesised by reacting 2-acetylpyridine with 4-nitrobenzaldehyde in base followed by ring closure with ammonium acetate according to methods outlined by E Constable et al (J Chem Soc Dalton Trans, 1992, 2947), followed

by reduction of the nitro group to the amine and quaternisation as described previously to form a compound described by formula I with R<sub>2</sub> and R<sub>3</sub> as hydrogen and R<sub>1</sub> as



10 Yet another general preparative method for a terpyridal group of general formula I is based on derivatising the structure below (Potts et al, JACS 1987, v109, 3961) by oxidising it to the carboxylic acid by e.g. the methodology outlined by Dodd et al (Synthesis (1993), V3, 295). Amination of the carboxylate is then carried out by standard procedures outlined in e.g. Chem Rev. (1981) V81, p589.



### Bipyridyls

The synthesis of a compound of general formula III with R<sub>8,9</sub> = NH<sub>2</sub> is described in (JACS 1958, V80, 2745) and (J Chem Soc Perkin Trans 2, 1996, 613) and a  
25 compound of formula III can be quaternised by the preceding methods outlined for terpyridyls.

### Phthalocyanines

Phthalocyanine dyes can be synthesised with amine nitrogen groups by e.g. Buchwald  
30 amination of halide precursors to produce outer-ring derivatives such as (J Org Chem 2000, V65, 1158), the amine groups of which are then quaternised.

### Tetra-aza-annulenes (TADAs)

TADAs of the general formula V can be derivatised at R1, R2, R3, and R4 by the general methods outlined above.

#### Dimers

- 5 The bipyridyl compounds tris(2,2'-bipyridyl - 4,4'-dicarboxylate) ruthenium (II) dichloride and tris(2,2'-bipyridyl) ruthenium (II) dichloride can be dimerised using pyrazine derivatives such as pyrazine, pyrimidine and 4,4'-bipyridyl linking ligands according to procedures detailed in (E A Seddon & K R Seddon, The Chemistry of Ruthenium, Elsevier, New York 1984, p 436).

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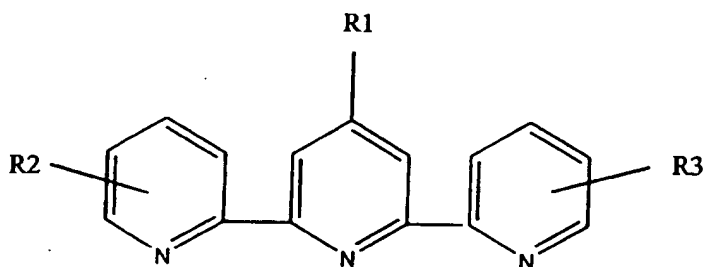
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**CLAIMS**

1. A composition comprising a photocatalyst, and a metal complex sensitiser comprising a ligand with a conjugated  $\pi$  system, which absorbs light substantially in the visible and/or the infrared region of the spectrum, effective to deposit a functional residue of said composition on a surface.
2. A composition as claimed in claim 1 wherein the atom within said metal complex sensitiser is any one or more of the following: ruthenium, platinum, palladium, iridium, rhodium, osmium, rhenium, iron, copper, titanium and zinc.
3. A composition as claimed in claim 1 or claim 2 wherein the metal complex sensitiser includes any one or more of the following: heterocyclic complexes which contain polypyridine, macrocyclic or phthalocyanine ligands and optionally other ligand types wherein at least one of the nitrogen groups is displaced by other donor groups
4. A composition as claimed in any preceding claim wherein the sensitising agent includes any one or more of the following: ruthenium II, III or IV or mixed oxidation state chelating complexes containing nitrogen donor atoms.
5. A composition as claimed in any one of claims 1 to 3 wherein the sensitising agent includes any one or more of the following: a ruthenium(II), (III), (IV) or a mixed oxidation state polypyridine complex.
6. A composition as claimed in claim 1 or claim 2 wherein said sensitising agent includes any one or more of the following groups: terpyridyl, bipyridyl groups, phthalocyanines, porphyrins, tetra-aza-annulenes, pyrazines, phenanthrolines and derivatives thereof and compounds with substantially similar nitrogen based ring systems.
7. A composition as claimed in claim 6 wherein said sensitising agent further includes any one or more of  $R_4N^+$  or  $R_4P^+$  groups wherein each R group may be the same or different and is any one or more of the following groups: hydrogen, halogen, amine, alkyl, aryl, arylalkyl, alkoxy, heterocyclic groups, or derivatives thereof,

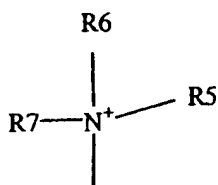
including acid and ester derivatives, any of which may be branched or unbranched, substituted or unsubstituted,.

8. A composition as claimed either of claim 6 and claim 7 wherein the sensitising agent includes a terpyridyl group of general formula I shown below:



(I)

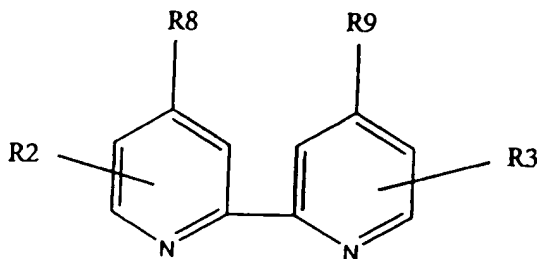
Where at least one of R1, R2 and R3 are positively charged groups which has the general formula II shown below:



(II)

Where R5-R7 are any one or more of the following groups: hydrogen, halogen, amine, alkyl, aryl, arylalkyl, alkoxy, heterocyclic groups, or derivatives thereof, including acid and ester derivatives, any of which may be branched or unbranched, substituted or unsubstituted,.

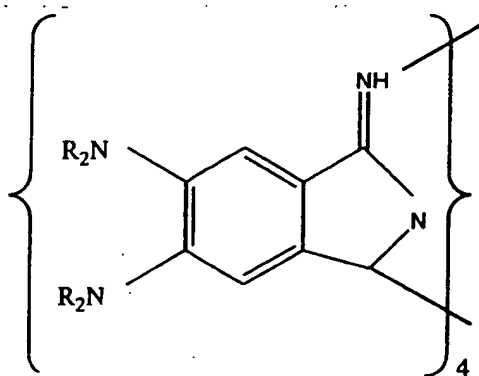
9. A composition as claimed in either claim 6 or claim 7 wherein the sensitising agent includes a bipyridyl group of the general formula (III) shown below:



## (III)

Where R8 and R9 can be the same or different and is any one or more of the following groups: hydrogen, halogen, amine, alkyl, aryl, arylalkyl, alkoxy, heterocyclic groups, or derivatives thereof, including acid and ester derivatives, any of which may be branched or unbranched, substituted or unsubstituted, R2 may be the same or different from R3 and is any one or more of the following groups: hydrogen, halogen, amine, alkyl, aryl, arylalkyl, alkoxy, heterocyclic groups, or derivatives thereof, including acid and ester derivatives, any of which may be branched or unbranched, substituted or unsubstituted,

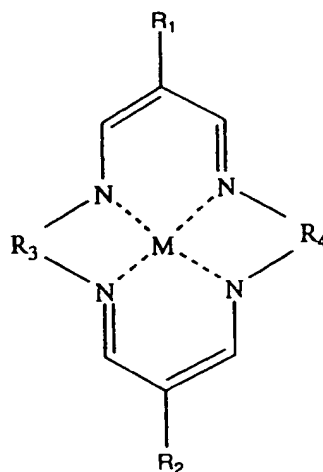
10. A composition as claimed in either claim 6 or claim 7 wherein the sensitising agent includes a phthalocyanine group of molecules and derivatives thereof having the general formula (IV) shown below:



## (IV)

Where each R group may be the same or different and are one or more the following groups: hydrogen, halogen, amine, alkyl, aryl, arylalkyl, alkoxy, heterocyclic groups, or derivatives thereof, including acid and ester derivatives, any of which may be branched or unbranched, substituted or unsubstituted,.

11. A composition as claimed in claim 6 or claim 7 wherein the sensitising agent includes a tetra-aza-annulene group and derivatives thereof and is described by the general formula shown (V) below:



(V)

15 Where R1-R4 may be the same or different and is any one or more of the following groups: hydrogen, halogen, amine, alkyl, aryl, arylalkyl, alkoxy, heterocyclic groups, or derivatives thereof, including acid and ester derivatives, any of which may be branched or unbranched, substituted or unsubstituted,

20 12. A composition as claimed in any preceding claim which is a liquid.

13. A composition any claimed in claim 12, wherein the light absorbing agent comprises no more than 50% w/v of said composition.

25 14. A composition as claimed in claim 12, wherein the sensitising agent comprises no more than 1% w/v of said composition.

15. A composition as claimed in any one of claims 12 to 14 which has a pH less than 7.

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16. A composition as claimed in any one of claims 12 to 14 which has a pH of 7 or more.

17. A composition as claimed in any one of claims 1 to 16 further comprising a doping agent.
18. A composition as claimed in claim 17 wherein the doping agent is any one or more of the following: platinum, palladium, cobalt, tungsten, chromium, silver, copper, nickel or iron, and/or compounds and complexes thereof.
19. A composition as claimed in any preceding claim further comprising a wetting agent which is any one or more of the following: Igepal® CA-520 [polyoxyethylene(5) isooctylphenyl ether], Igepal® CA-630 [(octylphenoxy)polyethoxyethanol], Igepal® CA-730 [polyoxyethylene(12) isooctylphenyl ether].
20. A sensitising agent which includes any one or more of the following groups : terpyridyl, bipyridyl groups, phthalocyanines, porphyrins, tetra-aza-annulenes, pyrazines, phenothralines and derivatives thereof and compounds with substantially similar nitrogen based ring systems.
21. A sensitising agent as claimed in claim 20 which further includes any one or more of  $R_4N^+$  or  $R_4P^+$  groups wherein each R group may be the same or different and is any one or more of the following groups: hydrogen, halogen, amine, alkyl, aryl, arylalkyl, alkoxy, heterocyclic groups, or derivatives thereof, including acid and ester derivatives, any of which may be branched or unbranched, substituted or unsubstituted,
22. A sensitising agent as claimed in claim 20 or 21 having one or more of the features of claims 8 to 11.
23. The use of a sensitising agent according to the present invention for the sensitisation of a light absorbing agent on a surface such that soils present on the surface are substantially broken down and/or the rate of accumulation of such soils on a surface is significantly diminished.



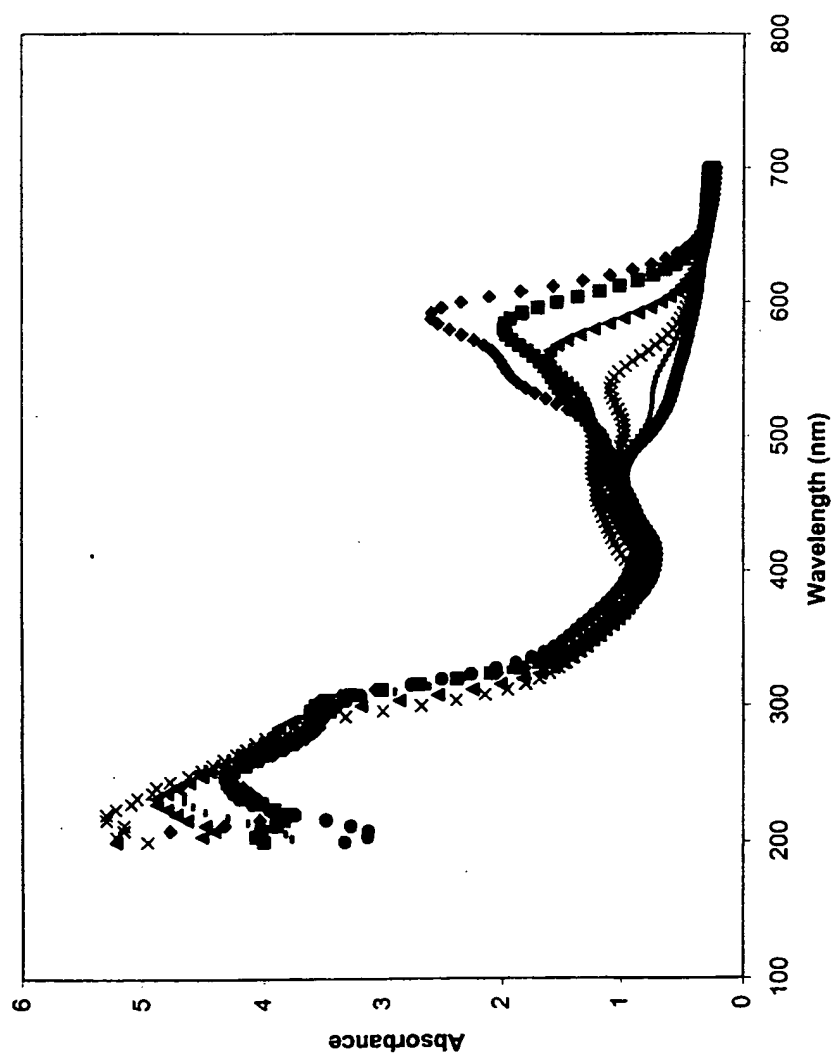


Figure 1

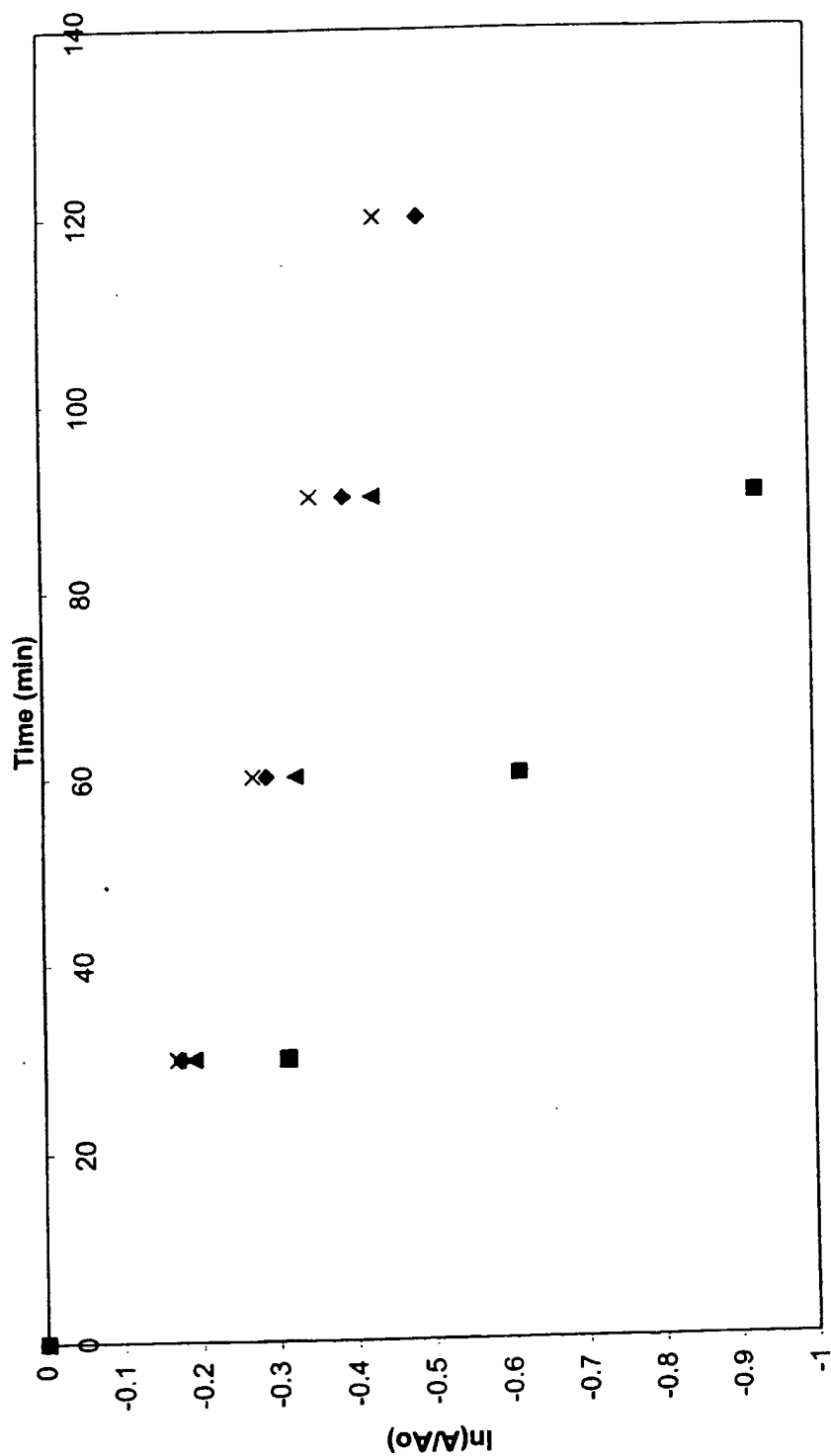


Figure 2

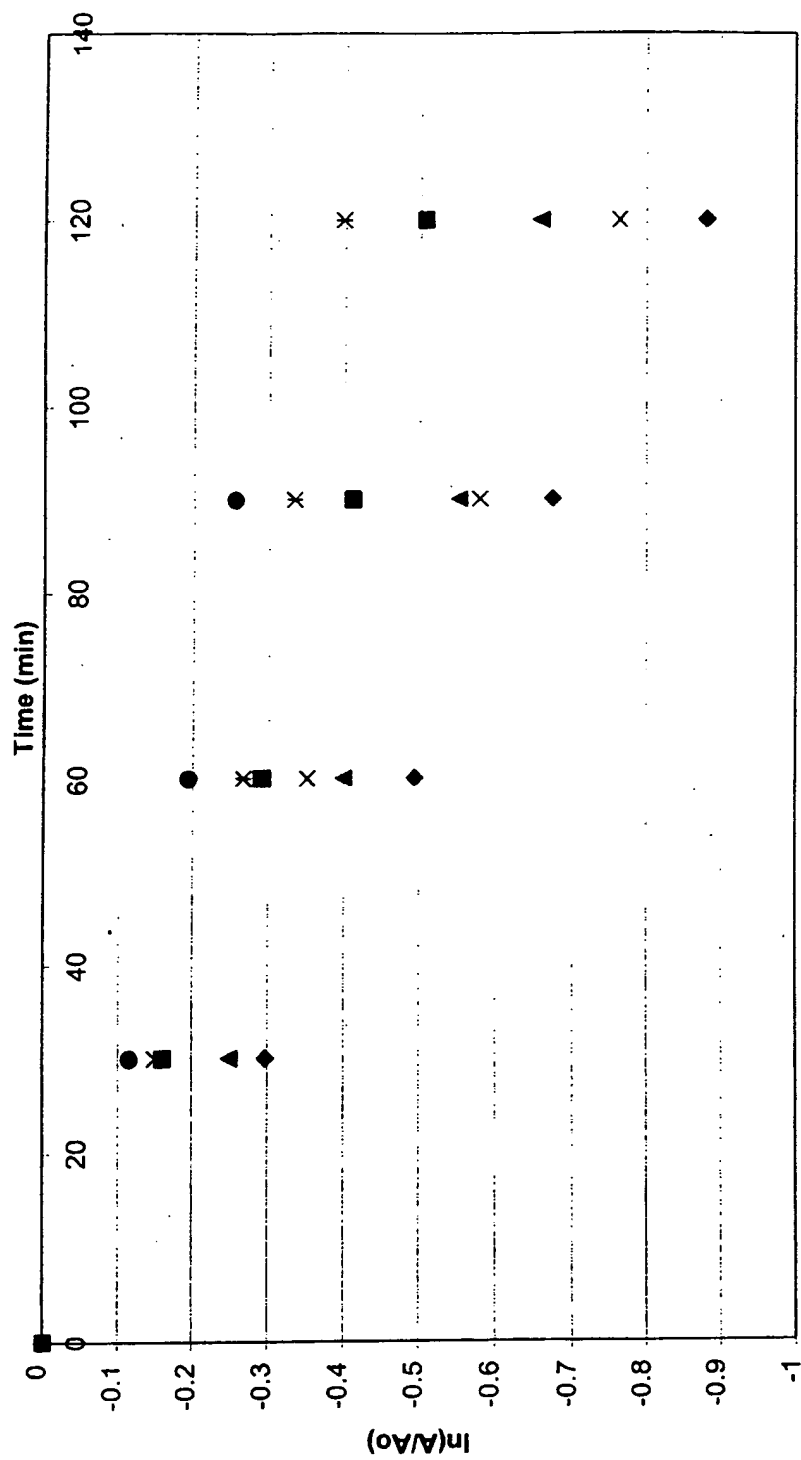


Figure 3

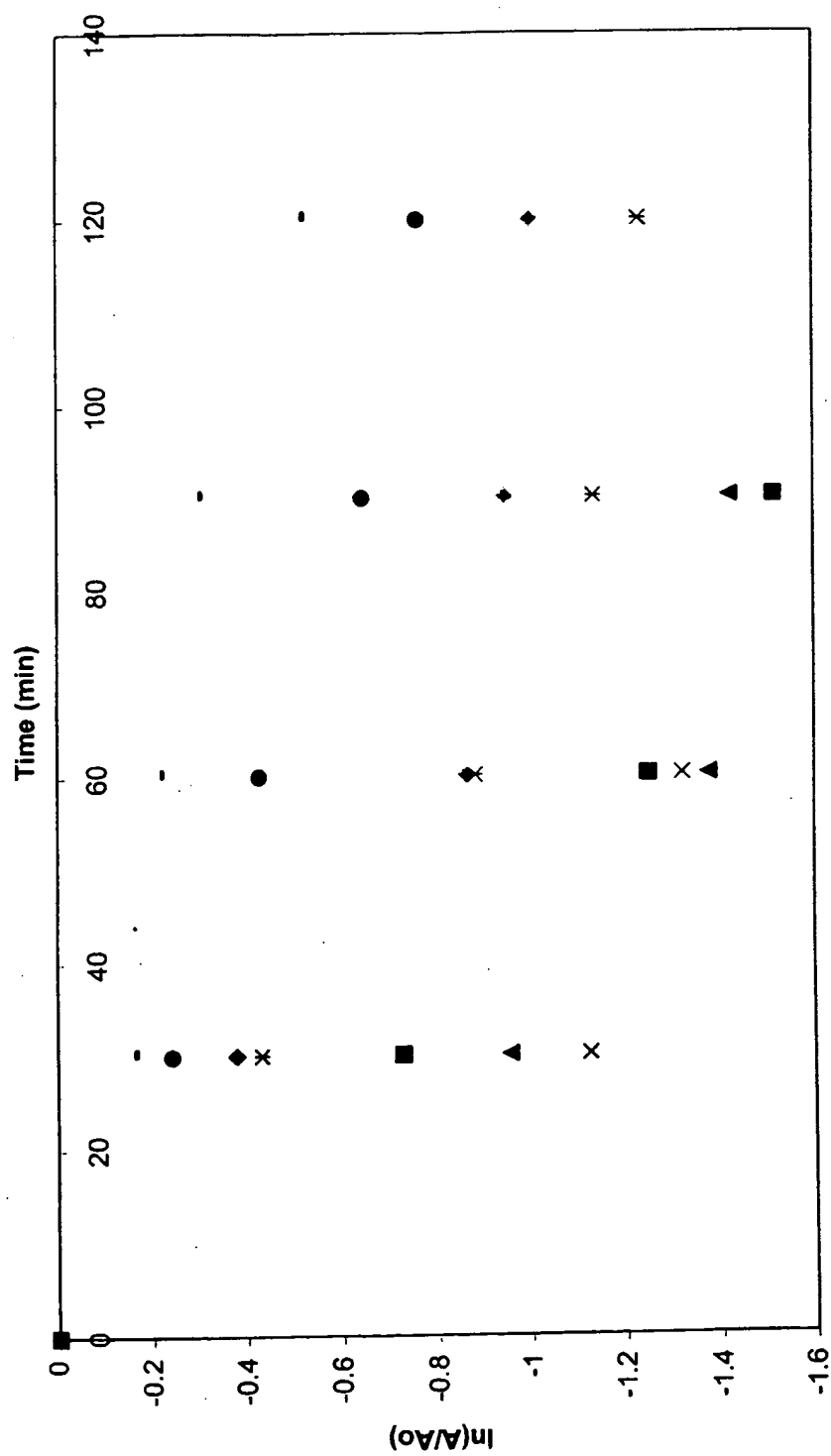


Figure 4

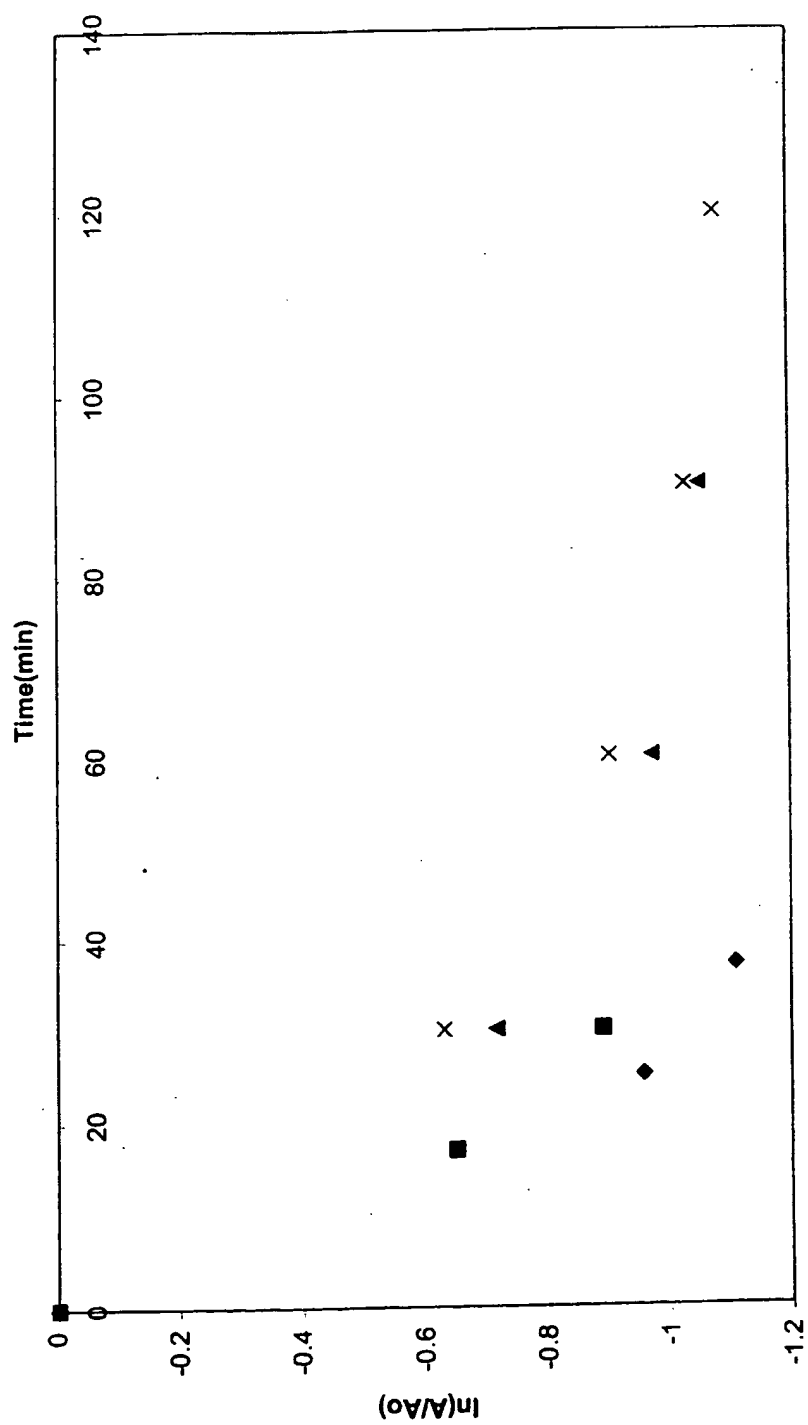


Figure 5

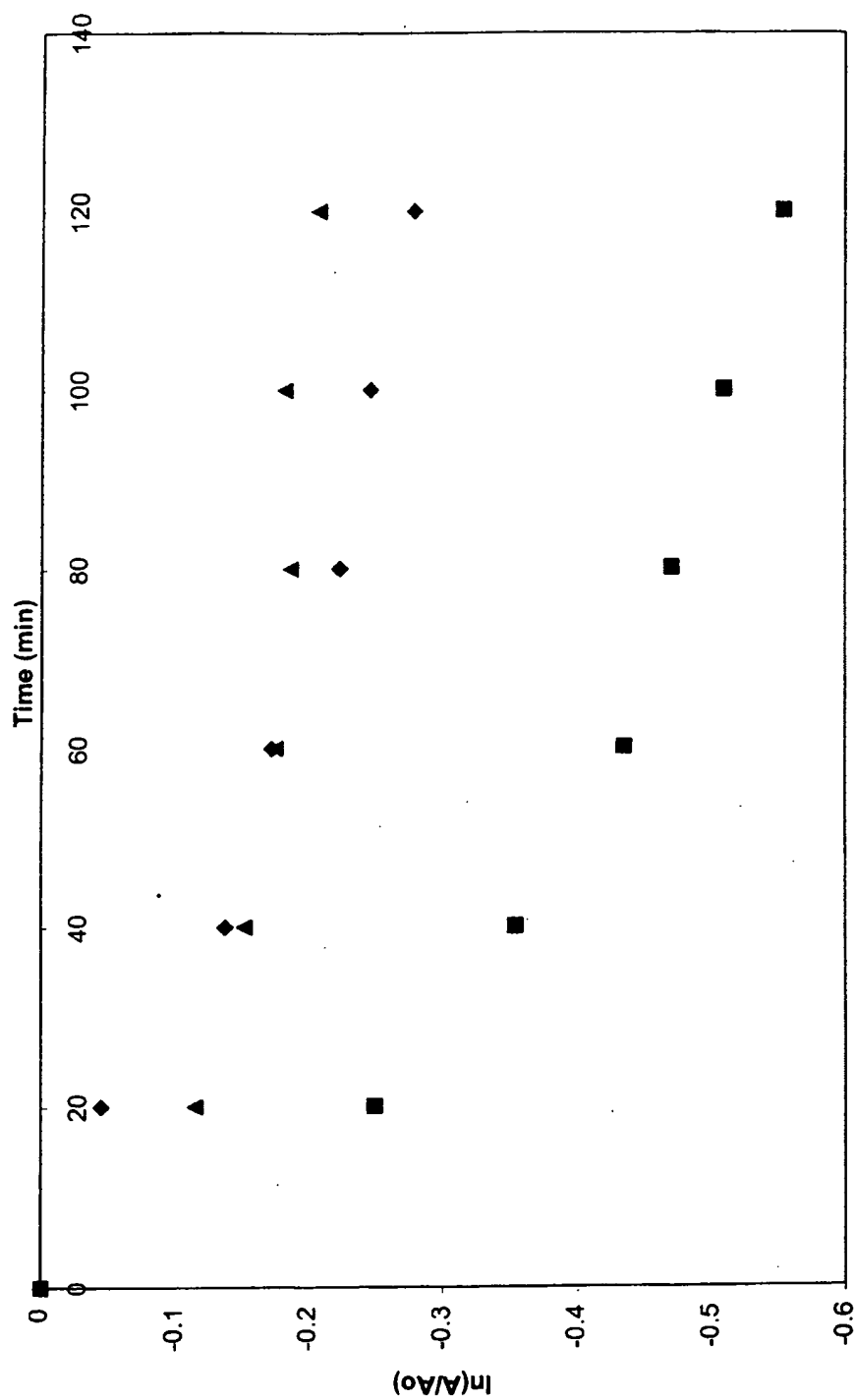


Figure 6

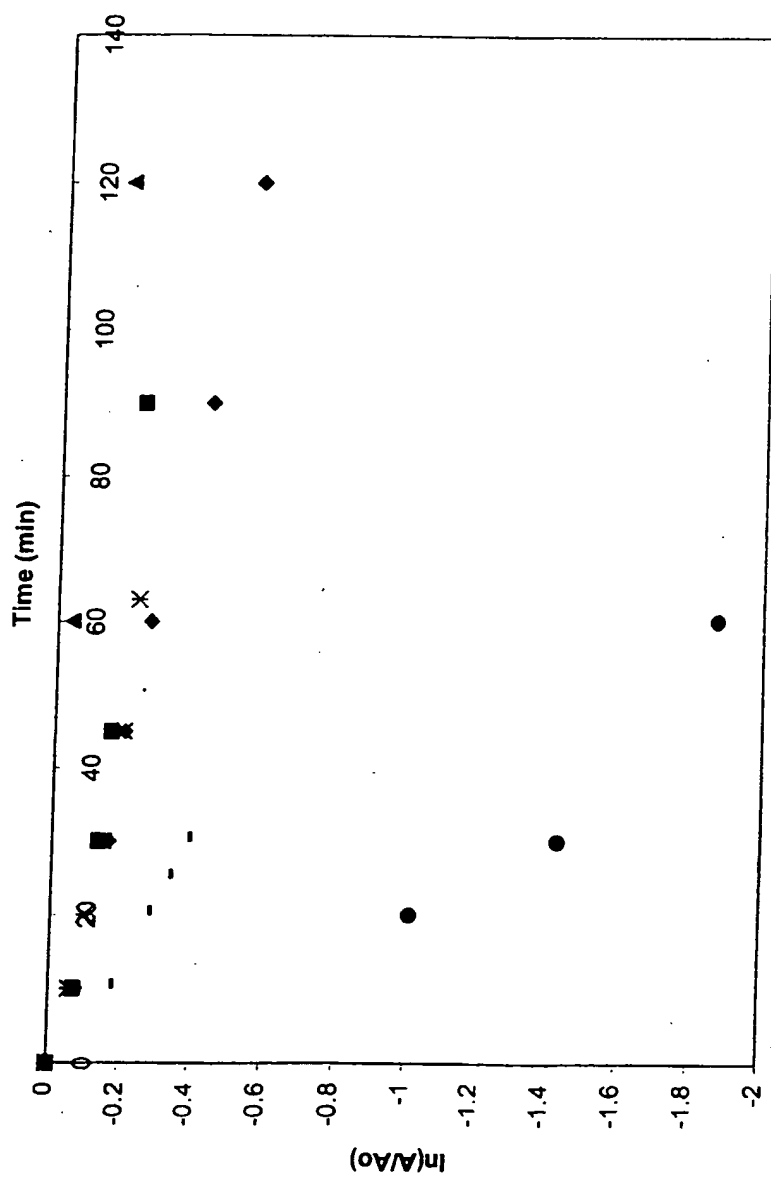


Figure 7